Amendments to the Specification:

Please replace paragragh starting at page 12, line 33 with the following amended paragraph:

There is no particular limitation to the material used for the substrate ~~5~~ 15. Materials such as alumina, zirconia, or the like may be used. Furthermore, if the substrate is electrically insulating, then it is easier to maintain insulation between the pair of separators 17. In addition, there is also no particular limitation on the material used for the seal glass 16. For example, the material used in general solid oxide fuel cells may be used.

Please replace paragragh starting at page 20, line 5 with the following amended paragraph:

First, a paste containing $LaMnO_3$ particles with an average particle diameter of 5 μm or less, $Ce_{0.9}Gd_{0.1}O_2$ particles with an average particle diameter of 5 μm or less, and carbon powder with an average particle diameter of 10 μm (manufactured by Nippon Carbon Co., Ltd.) was made up by mixing the materials noted above, and further adding propylene glycol and mixing. Next, a dry membrane with a thickness of 1 mm was formed by coating the paste onto a silica glass substrate using a printing method and heating the plate (at 120°C, for 60 minutes). Next, the $LaMnO_3$ particles and the $Ce_{0.9}Gd_{0.1}O_2$ particles were sintered by heating (at 1350°C, for 60 minutes) in an air atmosphere, and the cathode ($LaMnO_3$ / $Ce_{0.9}Gd_{0.1}O_2$ particle composite porous membrane, average pore diameter of 10 μm) with a thickness of 1 mm was formed by separating the cathode from the silica glass substrate. At this point the carbon powder was burnt off by oxidation. Continuing, the $Ce_{0.9}Gd_{0.1}O_2$ dense membrane (thickness 10 μm), which is the first solid ~~electrolyte~~ oxide, was formed by sputtering on the cathode that was formed, and a laminated body of the cathode and the first solid oxide were formed. At this point, a $Ce_{0.9}Gd_{0.1}O_2$ sintered body was used as the target of the sputtering.

Please replace paragraghs starting at page 27, line 15 with the following amended paragraphs:

Electrical power actually was generated with the samples thus fabricated, using hydrogen as the fuel and air as the oxidizing agent. Furthermore, when generating power, the power generation temperatures were set to be 400°C and 600°C, and the utilization factor of the anode

was 70% and that of the cathode was 40%. ~~FIG. 4~~ FIG. 6 shows the result of the power generation characteristics of sample 1 and sample 20 (comparative example).

As ~~FIG. 4~~ FIG. 6 shows, the result of sample 1 is power generation characteristics that are superior to that of sample 20, which is the comparative example. In particular, when the power generation temperature was 400°C, the degree of power output reduction was significantly less in sample 1 than the large reduction of the output of sample 20. It seems that the decrease of catalytic activity at low temperatures can be suppressed more in sample 1 than in sample 20.